

SHNEYDER, M.S., dotsent (Donetsk)

"Oxygen insufficiency in hypertension" by N.S.Zanozdra. Kaz.med.
zhur. no.3:90-91 My-Je '62. (MIRA 15:9)
(HYPERTENSION) (ANOXEMIA)
(ZANOZDRA, N.S.)

SHNEYDER, M.S., dotsent (Donetsk)

Disorders in the uniformity of pulmonary ventilation in anthracosis, pulmonary emphysema and pneumosclerosis of non-dust etiology. Klin.med. no.4:84-88 '62. (Klinika 15:5)

1. Iz kafedry propedevticheskoy terapii No.2 (zav. - prof. E.D. Borevskaya) Dontetskogo meditsinskogo instituta (dir. - dotsent A.M. Ganichkin) na baze 2-y gorodskoy bol'nitsy (glavnnyy vrach A.I. Solomakha).
(RESPIRATION) (LUNGS--DUST DISEASES) (EMPHYSEMA, PULMONARY)

VORONOV, Abram Solomonovich, prof.; KAMENETSKIY, S.I., red.; SHNEYDER,
M.S., red.; MAILYAN, S.L., red.; CHUCHUPAK, V.D., tekhn. red.

[Hospital therapy] Gospital'naia terapiia. Kiev, Gosmedizdat,
USSR, 1962. 522 p. (MIRA 16:2)

1. Zaveduyushchiy kafedroy gospital'noy terapii Donetskogo
meditsinskogo instituta (for Voronov).
(HOSPITAL THERAPY)

SHNEYDER, M.S., dotsent

"Functional examination of patients with silicosis" by I.M.
Ivanov, Kh.M.TSolov. Reviewed by M.S.Shneider. Sov.med.
(MIRA 16:4)
26 no.1:154-156 Ja '63.
(LUNGS--DUST DISEASES) (IVANOV, I.M.) (TSOLOV, Kh.M.)

SHNEYDER, M.S., dotsent

Massive hydrothorax as a complication of liver diseases. Kaz.
med. zhur. no.5:66-68 S-0163 (MIRA 16:12)

1. Klinika propedevticheskoy terapii (zav. dotsent M.I.
Frankfurt) Donetskogo meditsinskogo instituta.

RECORDED INFORMATION OF THE SUBJECTS OF THE LIST AND IN A
MATERIAL WHICH IS SUBMITTED, WILL NOT BE MAILED.
SUBMISSIONS 6269-011, JULY 1964
(WRI 1884)

1. UNCLASSIFIED AND UNMARKED MATERIAL.

SHINHYDER, M.Ye.; DENISOV, N.Ya., prof., red.; LEPESHINSKAYA, Ye.V., red.;
AKHTAMOV, S.N., tekhn. red.

[Technical terms in English, Russian, French, German, Swedish,
Portuguese, and Spanish used in soil mechanics and foundation
engineering] Slovar' tekhnicheskikh terminov po mekhanike
gruntov i fundamentostroenii na angliiskom, russkom, frantsuz-
skom, nemetskem, shvedskom, portugal'skom i ispanском языках.
Russkii tekst i ukazatel' M.H. Shneider, pod red. N.IA Denisova.
Moskva, Gos. izd-vo fiziko-matematicheskoi lit-ry, 1958. 139 p.
(MIRA 11:10)

1. International Society of Soil Mechanics and Foundation
Engineering.
(English language--Dictionaries--Polyglot)
(Soil mechanics--Dictionaries)

"APPROVED FOR RELEASE: 08/23/2000

CIA-RDP86-00513R001549810020-1

MYAKINA, Ye.B.; IVANOV, N.A.; SHNEYDER, N.M.

Carbohydrate and phosphorus metabolism in reactive states.

Probl. obshchei i sud. psikh. no.14:237-245 '63.

(MIRA 18;9)

APPROVED FOR RELEASE: 08/23/2000

CIA-RDP86-00513R001549810020-1"

FILIPPOVA, Mariya Filippovna, kand.geol.-miner.nauk; ARONOVA, S.M.; AFREMOVA, M.P.; GALAKTIONOVA, N.M.; GASSANOVA, I.G.; GIMPELEVICH, E.D.; KARASEV, M.S.; LYASHENKO, A.I.; MAYZEL', Z.L.; RATEYEV, M.A.; SOKOLOVA, L.I.; SOLOV'YEVA, N.S.; KHANIN, A.A.; SHISHENIHA, Ye.P.; ~~SHNEYDER, H.P.~~; BAKIROV, A.A., red.; VEBER, V.V., red.; DANOV, A.V., red.; DIKEN-SHTEYN, G.Kh., red.; MAKSIMOV, S.P., red.; POZNYSH, M.A., red.; SAIDOV, M.N., red.; SEMIKHATOVA, S.V., red.; TURKEL'TAUB, N.M., red.; UL'YANOV, A.V., red. [deceased]; KHALTURIN, D.S., red.; SHABAYEVA, Ye.A., red.; RAZINA, G.M., vedushchiy red.; GENNAD'YEVA, I.M., tekhn. red.

[Devonian deposits in the central provinces of the Russian Platform]

Devonskie otlozheniya tsentral'nykh oblastei Russkoi platformy.

Pod red. M.F. Filippovoi. Leningrad, Gos. nauchno-tekhn. izd-vo neft.

i gorno-toplivnoi lit-ry, 1958. 404 p. (MIRA 11:4)

(Russian Platform--Geology, Stratigraphic)

VITOSHINSKAYA, M.I., bibliograf; GEKKER, I.F., bibliograf; SHNEYDER, R.A.,
bibliograf; SOLOV'YEV, S.P., doktor geologicheskikh nauk, redaktor;
KULIKOV, M.V., kandidat biologicheskikh nauk, redaktor; PERLIN, S.S.,
redaktor izdatel'stva; GUROVA, O.A., tekhnicheskiy redaktor

[Geological literature of the U.S.S.R.; a bibliographical annual for
1951] Geologicheskaya literatura SSSR; bibliograficheskii ezhegodnik
za 1951 g. Moskva, Gos. nauchno-tekhn. izd-vo lit-ry po geol. i
okhrane nedr, 1956. 146 p. (MLRA 10:2)

1. Moscow. Vsesoyuznaya geologicheskaya biblioteka. 2. Vsesoyuznaya
geologicheskaya biblioteka Vsesoyuznogo Nauchno-issledovatel'skogo
geologicheskogo instituta Ministerstva geologii (for Vitoshinskaya,
Gekker, Shneyder, Solov'yev, Kulikov)
(Bibliography--Geology)

VITOSHINSKAYA, M.I., bibliograf; GEKKER, I.F., bibliograf; SHNEYDER, R.A.,
bibliograf; GLAZKOVSKAYA, Ye.A.; KLYASHTORNYY, S.G.; SOLOV'IEV,
S.P., doktor geologo-mineral.nauk, red.; KULIKOV, M.V., kand.
biolog.nauk, red.; FERLIN, S.S., red.izd-va; GUROVA, O.A.,
tekhn.red.

[Geological literature of the U.S.S.R.; a bibliographical year-
book for 1954] Geologicheskaya literatura SSSR; bibliograficheskii
ezhegodnik za 1954 g. Moskva, Gos. nauchno-tekhn.izd-vo lit-ry po
geol. i okhrane nedr, 1957. 185 p. (MIRA 12:1)

1. Moscow. Vsesoiuznaya geologicheskaya biblioteka.
(Bibliography--Geology)

VITOSHINSKAYA, M.I., bibliograf; GEKKER, I.F., bibliograf; SHNEYDER,
R.A., bibliograf; GLAZKOVSKAYA, Ye.A., bibliograf; KLYASHTORNYY,
S.G., bibliograf; SOLOV'YEV, S.P., doktor geologo-mineralog. nauk,
red.; KULIKOV, M.V., kand.biolog.nauk, red.; IVANOVA, A.G., tekhn.
red.

[Geological literature in the U.S.S.R.; bibliographical year-book
for 1955] Geologicheskaya literatura SSSR; bibliograficheskii
ezhegodnik za 1955 g. Moskva, Gos.nauchno-tekhn.izd-vo lit-ry po
geologii i okhrane nedr, 1959. 333 p. (MIRA 12:11)

1. Moscow. Vsesoyuznaya geologicheskaya biblioteka. 2. Vseso-
yuznaya geologicheskaya biblioteka Vsesoyuznogo geologicheskogo
nauchno-issledovatel'skogo instituta.(for Vitoshinskaya, Gekker,
Shneyder, Glazkovskaya, Klyashtornyy).
(Bibliography--Geology)

SHNEYDER, R.G.; SHCHERBO, I.A.

Jupiter in 1948. BiulVAGO no.11:26-30 '52.

(MLRA 6:6)

1. Moskovskoye (fieleniye Vsesoyuznogo astronomo-geodezicheskogo obshchestva, otdel planet i Luny.
(Jupiter (Planet))

SHNEYDER, R.G.

Eclipsing variable IS Cassiopeiae. Per.zvezdy 12 no.4:312-314
Je '58. (MIRA 13:4)

1. Gosudarstvennyy astronomicheskij institut im.P.K.Shternberga.
(Stars, Variable)

SHNEYDER, R.G.

Long-period variable RT Cephei. Per.zvezdy 12 no.4:315-317 Je
'58. (MIRA 13:4)

1. Astronomicheskiy sovet AN SSSR.
(Stars, Variable)

SHNEYDER, R. I.

PA 10/49T43

USSR/Engineering
Fuel Consumption
Gas, Natural

Jul 48

"Familiarization of Gas Burning by Petroleum Industry Undertakings," R. I. Shneyder, OrgEnergoNeft,
 $4\frac{1}{2}$ pp

"Energet Byul" No 7

Describes properties of natural gas. Explains how it should be used for fuel. Discusses experience.

10/10/86

SHNEYDER, R.I.

Ripple current supply of testing circuits for electric measuring
instruments. Izm.tekh. no.6:38-40 Je '65.

(MIRA 18:8)

IVANOV, Boris Nikolayevich; TKALIN, Ivan Mikhaylovich; SOLNTSEV, Vyacheslav Aleksandrovich; SHTRUM, Viktor L'vovich; SHNEYDER, Roman Izraylevich; MAYANSKIY, Iosif Isaakovich; BORISOVA, Volya Petrovna; ARUTYUNOV, V.O., retsentent; BLEKHSHTEYN, L.I., red.; SOBOLEVA, Ye.M., tekhn.red.

[Technology of the manufacture of electric instruments] Tekhnologiya elektropriborostroenia. Moskva, Gos.energ.izd-vo, 1959.
590 p. (MIRA 13:4)

(Electric apparatus and appliances)

SHNEYDER, R.M.

Improving the utilization of working capital with established norms
in railroads. Trudy MIIT no.119:158-185 '59. (MIRA 12:11)
(Railroads--Finance)

SHNEYDER, R.M.

Bank credit to finance fundamental railroad work. Trudy MIIT no.119:
186-195 '59. (MIRA 12:11)
(Railroads--Finance) (Credit)

SHNEY DER, R.M.

Amortization deductions made by the railroads of capitalist
countries. Zhel.dor.transp. 42 no.7:86-90 J1 '60.
(MIRA 13:7)

(Railroads--Finance)

"APPROVED FOR RELEASE: 08/23/2000

CIA-RDP86-00513R001549810020-1

STRCGANOV, G.B.; SHNEYDER, S.A.; ENTIN, L.Kh.

Device for checking the gas saturation of cast aluminum alloys.
Zav. lab. 31 no.8:1030 '65. (MIRA 18:9)

APPROVED FOR RELEASE: 08/23/2000

CIA-RDP86-00513R001549810020-1"

KHARITONOVА, K. K.

Repair of cranial defects with horn plates. Vopr. neirokhir. 16 no. 4:
37-41 July-Aug 1952. (CLML 23:3)

1. Senior Scientific Associate. 2. Of the Clinical Department (Head --
Prof. S. L. Shneyder), Novosibirsk Scientific-Research Institute of
Orthopedics and Restorative Surgery (Director -- Docent D. P. Matelkin).

SERE BROV, M.A. (Novosibirsk, Kamennoye shosse, d.5, kv.12)

Anatomical surgical principles of ligation of the gluteal arteries.
Vest.khir. 77 no.9:54-60 S '56. (MIRA 9:11)

1. Iz Novosibirskskogo nauchno-issledovatel'skogo instituta vosstano-
vitel'noy khirurgii i ortopedii (dir. - dotsent D.P.Metelkin,
rukoved. raboty prof. S.L.Shneyder)
(BUTTOCKS, blood supply
gluteal arteries, surg. ligation, technic & surg. anatomy)

VI'TSYN, B.A.

Treatment of sacro- and coccygeorectal fecal fistulas. Khirurgia,
Moskva no. 2:64-68 Feb 1953. (CLML 24:2)

1. Candidate Medical Sciences. 2. Of the Clinic of Hospital Surgery
(Director -- Prof. S. L. Shneyder), Novosibirsk Medical Institute,
and of the Clinic of Restorative Surgery, Novosibirsk Scientific-Re-
search Institute for Restorative Surgery and Orthopedics.

SHNEYDER, S. L., Prof.

Head, Clinic of hospital surgery Novosibirsk medical institute

"Unusual localization of fecal fistula with atypical course," by B. A. Vitsyn, Vest. khir. 72 no. 4 Jl-Ag 1952.

SKUMBIN, M.K.; SOLONININ,A.V.;SHNEYDER, T.M.; RYASHKO, B.V.; GAVRYUSHIN, N.M.; KHARLANOVICH, I.V.

Complex technology for train and freight operations in a division.
Zhel. dor. transp. 46 no.8:14-21 Ag '64.

(MIRA 17:11)

1. Nachal'nik Permskogo otdeleniya Sverdlovskoy dorogi (for Skumbin).
2. Zamestitel' nachal'nika Permskogo otdeleniya Sverdlovskoy dorogi (for Soloninin).
3. Glavnyy inzh. Permskogo otdeleniya Sverdlovskoy dorogi (for Shneyder).
4. Nachal'nik otdela dvizheniya Permskogo otdeleniya Sverdlovskoy dorogi (for Ryashko).
5. Zamestiteli nachal'nika otdela dvizheniya Permskogo otdeleniya Sverdlovskoy dorogi (for Gavryushin, Kharlanovich).

SHNEYDER, V., inzh.

New coal-mining equipment. NTO no.10:62 O '59.

(MIRA 13:2)

1.Chlen Nauchno-tekhnicheskogo obshchestva Gornoye, g.Anzhero-Sudzhensk,
Kemerovskoy oblasti.

(Anzhero-Sugzhensk--Coal mining machinery)

SHNEYDER, V., inzh.

Visiting our counterparts in the Donets Basin. NTO 2 no. 5:40 My
'61. (MIRA 14:5)

1. Shakhta 9-15 g. Anzhero-Sudzhensk, Kemerovskoy oblasti, chlen
delegatsii gornyyakov Kuzbassa.

(Donets Basin--Coal mines and mining)
(Kuznetsk Basin--Coal mines and mining)

MAKHLIN, Z., inzh. (Leningrad); SHNEYDER, V. (g. Anzhero-Sudzhensk,
Kemerovskoy oblasti); IVANNIKOV, V., inzh. (Novosibirsk);
PEKELIS, G., inzh. (Leningrad); KIRYUSHIN, N., inzh. (Krasnodar)

Suggested, created, introduced. Izobr. i rats. no. 7:20-21 J1 '61.
(MIRA 14:6)

1. Zamestitel' predsedatelya soveta Vsesoyuznogo obshchestva
izobretateley i ratsionalizatorov obogatitel'noy fabriki 9-15
(for Shneyder).

(Technological innovations)

KHOKHOLEV, K. [Khokholiev, K.], inzh.; SAVIN, M., inzh.; SHNEYDER, V., inzh.

Floor slabs made of slag foamed concrete. Bud. mat. i konstr. 4 no.1:
22-24 Ja-F '62. (MIRA 15:7)
(Floors, Concrete) (Lightweight concrete)

FASON, E. (Bukharest); SHVARTS, B. (Bukharest); SHNEYDER, V. (Bukharest)

Blood penetration into the cerebral ventricles. Zhur. nevr.
i psikh. 64 no.2:212-218 '64. (MIRA 17:5)

SHNEYDER, V. A.

USSR/Engineering - Construction work

Card 1/1 : Pub. 70 - 2/11

Authors : Shneyder, V. A. Engineer

Title : About output increase of wheel-type excavators

Periodical : Mekh. stroi. 4, 6-9, Apr 1954

Abstract : A 5-point program for increasing the work-output of wheel-type excavators D-183, D-188, D-189 and D-213 (all Soviet made vehicles), is presented. The program also lists the economical gains to be derived from increasing the output of ground excavators. Tables; illustrations; drawings.

Institution :

Submitted :

SUMYANOV, V. A.

"Reference List of Candidates of Technical Sciences Dissertations on 'The Efficient Shape and Size for the Bowl of a Wheel-Mounted Scraper'." Cand Tech Sci, Moscow Inst Municipal Construction Engineers of the Moscow City Executive Committee, Moscow, 1955. (KL, No 8, Feb 55)

Su: Sum. No. 631, 26 Aug 55-Survey of Scientific and Technical Dissertations Defended at USSR Higher Educational Institutions (14)

YEL'CHIN, B.; SHNEYDER, V., kandidat ekonomicheskikh nauk.

Prospective development in the production of roofing materials.
Stroi.mat.izdel. i konstr. 1 no.12:21-22 D '55. (MLRA 9:?)
(Roofing)

BERKOVICH, T., kandidat tekhnicheskikh nauk; RABINOV, I., kandidat tekhnicheskikh nauk; SOLNTSEVA, V., kandidat tekhnicheskikh nauk; SMIRNOV, N., doktor geologo-meneral'nyy nauk; SHNEYDER, V., kandidat ekonomicheskikh nauk.

Making roof slate and asbestos pipes using a sand cement base.
Stroi.mat., izdel.i konstr. 1 no.11:4-6 N '55. (MLRA 9:5)
(Roofing) (Asbestos cement)

SHNEYDER, V.A.

Side guard for buckets of the D-147 and D-222 scrapers. Rats.
izobr. predl. v stroi. no.112:15-16 '55. (MLRA 9:6)
(Scrapers)

SHNEYDER, V.A., kand.tekhn.nauk

Hinged bulldozer end blades. Mekh.stroi. 14 no.6:12-14 Je '57.
(MIRA 10:11)
(Bulldozers)

SHNEYDER, V.A., kand. tekhn. nauk

Increasing productivity of graders in leveling. Transp. stroi.
8 no. 5:20-22 My '58. (MIRA 11:7)
(Road machinery)

SOKOLOV, K.N.; YEVSTAFYEV, S.V.; ROSTOTSKIY, V.K.; GRECHIN, N.K.; STANKOVSKIY, A.P.; BAUMAN, V.A.; BERKMAN, I.L.; BORODACHEV, I.P.; BOYKO, A.G.; VALUTSKIY, I.I.; VATSSLAVSKAYA, L.Ya.; VOL'FSION, A.V.; DOMBROVSKIY, N.G.; YEGNUS, M.Ya.; YEFREMENKO, V.P.; ZIMIN, P.A.; IVANOV, V.A.; KOZLOVSKIY, A.A.; KOSTIN, M.I.; KRIMERMAN, M.N.; LINEVA, M.S.; MERENKOV, A.S.; MIROPOL'SKAYA, N.K.; PETROV, G.D.; REBROV, A.S.; ROGOVSKIY, L.V.; SMIRNOV, G.Ya.; SHAFRANSKIY, V.N.; SHIMANOVICH, S.V.; SHNEYDER, V.A.

Evgenii Richardovich Peters; obituary; Melch. stroi. 15 no.1:3 of cover
Ja '58. (MIRA 11:1)

(Peters, Evgenii Richardovich, 1892-1957)

SHNEYDER, Viktor Aleksandrovich; MANUYLOV, Yu.G., nauchnyy red.;
SIDEL'NIKOVA, L.A., red.; KOZLOVSKAYA, M.D., tekhn.red.
PERSON, M.N., tekhn.red.

[Scrapers, bulldozers, graders] Skrepery, bul'dozery,
greidery. Moskva, Proftekhnizdat, 1961. 235 p. (MIRA 15:5)
(Earthmoving machinery)

SHNEYDER, Viktor Aleksandrovich; MANUYLOV, Yu.G., nauchn. red.;
NAZARENKO, M.I., red.

[Scrapers, bulldozers, and graders] Skrepery, bul'dozery,
greidery. Izd.2., ispr. i dop. Moskva, Vysshiaia shkola,
1964. 269 p. (MIRA 18:3)

SELIVANOVA, N.M.; SHNEYDER, V.A.

Refractive index and refraction of selenates. Zhur. fiz. khim.
38 no.7:1822-1824 J1 '64. (MIRA 18:3)

I. Khimiko-tehnologicheskiy institut imeni Mendeleyeva.

Shneyder, V.A.

32-11-9/60

AUTHORS: Persiantseva, V.P., Shneyder, V.A.

TITLE: Colorimetric Determination of the 2.6-, 2.7-Naphthalene Disulphide Acid Content in a Nickel Electrolyte (Kolorimetricheskoye opredeleniye 2.6-, 2.7-naftalindisul'fokisloty v nikel'evom elektrolite)

PERIODICAL: Zavodskaya Laboratoriya, 1957, Vol. 23, Nr 11, pp. 1298-1298 (USSR)

ABSTRACT: The acid mentioned is characterized by the fact that it causes the "phenomenon of mould" in the nickel electrolyte. For the purpose of determining it, it is recommended to carry out separation of nickel from the electrolyte chromatographically by the application of the cationite "SDV-3". For this purpose an adsorber was mounted in a cylindrical vessel (filter) of 15 mm diameter up to a height of 200 mm. The electrolyte was made to pass through the adsorber with a velocity of 100 ml per 20 min. After nickel adsorption was completed the filter was rinsed out while the same velocity was maintained for 10 minutes. The analysis itself was then carried out: 5 ml of the electrolyte was dissolved in distilled water of up to 20-25 ml total volume and sent through the adsorption filters. The filters were rinsed out with distilled water; the solution together with the rinsing water was introduced into a platinum vessel, where 10 ml

Card 1/2

AUTHORS: Mel'nikov, S. N., Shneyder, V. A. SCV/156-58-2-4/48

TITLE: Concerning Some Physicochemical Properties of Barium-Selenate
(O nekotorykh fiziko-khimicheskikh svoystvakh selenata bariya)

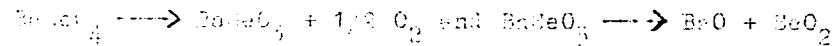
PERIODICAL: Nauchnyye doklady vysshey shkoly, Khimiya i khimicheskaya
tekhnologiya, 1958, Nr 2, pp. 216-220 (USSR)

ABSTRACT: Salenates of alkaline earth are more and more used in industry. However, the properties mentioned in the title, are almost unknown in pertinent publications (Refs 1-5). This statement is illustrated by some examples (Refs 6-9). The authors determined the solubility of barium selenate polarographically in water of temperatures ranging from 15-95° and studied the properties of crystalline salt in the course of a heat treatment. The values of solubility are shown in table 1. The composition and crystalline structure of the bottom phase undergoes no modification between 15 and 100° and is formed at a stir-up of the mentioned salt in water. The thermal decomposition process sets in at about 350°. It proceeds in solid state with the absorption of heat. Barium selenate does not melt; its by-product barium selenite, however, does. The melting temperature of the latter amounts to about 1285°. The thermal decompo-

UDC 156-53-2-4/49

Concerning Some Physico-chemical Properties of Barium-Selenate

reaction process of barium selenate proceeds according to the following equation:



Professor A. F. Kapustinskiy participated in the critical discussion of this project. There are 2 figures, 3 tables, and 10 references, 4 of which are Soviet.

ASSOCIATE: Sof'ya na rogo-nicheskoy khimii Moskovskogo khimiko-tehnologicheskogo instituta im. D. I. Mendeleyeva
(Chair of Inorganic Chemistry of the Chemical and Technical Institute named D. I. Mendeleyev, Moscow)

SUBMITTED: September 31, 1957

Card 2/2

AUTHORS: Selivanova, N.M., Shneyder, V.A. SOV/63-3-6-37/43

TITLE: The Thermodynamic Properties of Selenate Plaster ($\text{CaSeO}_4 \cdot 2\text{H}_2\text{O}$)
(O termodinamicheskikh svoystvakh selenato-gipsa) ($\text{CaSeO}_4 \cdot 2\text{H}_2\text{O}$)

PERIODICAL: Khimicheskaya nauka i promyshlennost', 1958, Vol III, Nr 6,
pp 834-835 (USSR)

ABSTRACT: The formation heat of selenate plaster has been determined as
9.263 kcal/mole, the change of the free energy as -335.18
kcal/mole, the absolute entropy as 36.08 entropic units.
There are 3 references, 2 of which are Soviet and 1 American.

ASSOCIATION: Moskovskiy khimiko-tehnologicheskiy institut imeni D.I. Men-
deleyeva (Moscow Chemical-Technological Institute imeni D.I.
Mendeleev)

SUBMITTED: May 14, 1958

Card 1/1

78-3-6-4/30

AUTHORS: Selivanova, N. M., Shneyder, V. A., Zubova, G. A.

TITLE: On the Thermal Decomposition of the Selenates of Strontium,
Barium and Lead (O termicheskem razlozhenii selenatov
strontsiya, bariya i svintsa)

PERIODICAL: Zhurnal Neorganicheskoy Khimii, 1958, Vol. 3, Nr 6,
pp. 1295 - 1303 (USSR)

ABSTRACT: The thermograms and the cooling curves of the selenates of strontium, barium and lead were investigated in order to explain the effects occurring in these curves. The thermographic analyses of strontium and barium selenate were performed within temperature ranges of from 100-1300°C. On this occasion three effects take place: For strontium selenate: at 525°C(exothermic), at 835°C(endothermic), and at 115°C (endothermic). For barium selenate: at 630°C(exothermic), at 900°C(endothermic), and at 1285°C(endothermic). The cooling curves of strontium and barium selenate do not agree with the heating curves of the two compounds, i. e. both processes are

Card 1/2

78-3-6-4/30

On the Thermal Decomposition of the Selenates of Strontium, Barium and Lead

not reversible. It can be seen from the thermograms that also selenium is oxidized by Se^{4+} in Se^{6+} . It was shown that the selenates of strontium, barium and lead when heated, pass into selenite under the release of oxygen. The chemical analyses of the final products in the thermal analyses were confirmed by x-ray analysis. The thermographic analysis of lead selenate showed that at 680°C and 930°C thermal effects take place and that beginning with 930°C this compound melts. At 1000°C the lead selenite formed at 700°C passes into lead oxide. It was found that strontium, barium and lead selenate are thermally more unstable than the corresponding sulfates. There are 7 figures, 6 tables and 33 references, 9 of which are Soviet.

SUBMITTED: May 6, 1957

AVAILABLE: Library of Congress

Card 2/2 1. Strontium selenates--Thermal analysis 2. Lead selenates--Thermal analysis 3. Barium selenates--Thermal analysis

SOV/156-53-4-13/49

AUTHORS:

Selivanova, N. M., Shneyder, V. A.

TITLE:

On Some Physico-Chemical Properties of Calcium Selenate
(O nekotorykh fiziko-khimicheskikh svoystvakh selenata
kal'tsiya)

PERIODICAL:

Nauchnyye doklady vysshyey shkoly. Khimiya i khimicheskaya
tekhnologiya, 1958, Nr 4, pp 664-666 (USSR)

ABSTRACT:

The solubility of $\text{CaSeO}_4 \cdot 2\text{H}_2\text{O}$ in water within the temperature range of from (-2.6°) to $(+101.8^\circ)$ was investigated, and the behavior of this salt on heating was determined. The chemical analyses and the microscopic and radioscopic investigations have shown that in the temperature range to be investigated only the compound $\text{CaSeO}_4 \cdot 2\text{H}_2\text{O}$ is present. The heating and cooling curve of calcium selenate for temperatures of from 50° to 1000°C were plotted. The heating curve of calcium selenate is up to 600° similar to the heating curve of gypsum. At 140°C the weight loss amounts to 13.46%, at 200°C to 16.40%. On the heating curve an endothermal effect occurs at 698°C , which indicates the decomposition of CaSeO_4 . The decomposition

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On Some Physico-Chemical Properties of Calcium Selenate

leads to the formation of selenite (CaSeO_3) and is finished at 800°C . The radiograms of the decomposition salts taken prove the presence of calcium selenite. The melting point of calcium selenite is at 920°C . The comparison of the melting temperature with the decomposition temperature of the selenates and sulfates of calcium shows that calcium selenate is thermally less stable than calcium sulfate. There are 3 figures, 1 table, and 8 references, 6 of which are Soviet.

ASSOCIATION: Kafedra obshchey i neorganicheskoy khimii Moskovskogo khimiko-tehnologicheskogo instituta im. D. I. Mendeleyeva (Chair of General and Inorganic Chemistry at the Moscow Chemo-Technological Institute imeni D. I. Mendeleyev)

SUBMITTED: May 9, 1958

Card 2/2

SHELYAKIN, V.I., Cand Chem Sci — (diss) "Thermal properties of selenides of metals of the second group (basic subgroup) of D.I. Mendeleev's periodic system of elements." Mos, 1959. 11 pp (Min of Higher Education USSR. Mos Order of Lenin Chemical Technological Inst im. D.I. Mendeleev. Chair of General and Inorganic Chemistry).

10 copies (KL, 39-59, 101)

17

5(1,4)

SOV/153-2-4-1/32

AUTHORS: Selivanova, N. M., Shneyder, V. A.

TITLE: Physico-chemical Properties of Selenates. V. Formation Heat of Calcium Selenate From Elements

PERIODICAL: Izvestiya vysshikh uchebnykh zavedeniy. Khimiya i khimicheskaya tekhnologiya, 1959, Vol 2, Nr 4, pp 475 - 479 (USSR)

ABSTRACT: The salt mentioned in the subtitle is very close to calcium sulfate on account of its properties. Although the heat effects of many reactions with the latter participating were described (Refs 1-5), investigations with regard to calcium selenate are missing. The paper under review aims at determining the formation heats of the crystalline salts $\text{CaSeO}_4 \cdot 2\text{H}_2\text{O}$ and CaSeO_4 .

Table 1 shows the determination results of the heat of the interacting reaction of the former salt with an aqueous silver nitrate solution. An average value of 9.26 ± 0.240 kcal/mol is assumed by the authors. The precipitate of silver selenate formed in the calorimeter was chemically analyzed. The results are in perfect agreement with data on a salt synthesized earlier. The same is true of the interplanar spacings computed on account of the

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Physico-chemical Properties of Selenates. V. Formation SOV/153-2-4-1/32
Heat of Calcium Selenate From Elements

Debye-Scherrer radiograph (Table 2). The formation heat of crystalline calcium sulfate (Ref 14) is shown in comparison, and the difference between the formation heats of the latter and the selenate (=70.29 kcal/mol) is determined. Table 3 shows the dissociation heats of the two selenates under discussion at 25°. A. A. Mayer determined the refractive indices. The hydration heat of CaSeO_4 ($Q_{\text{hydration}}$) was -6.02 ± 0.10 kcal/mol as compared to the hydration heat of -6.990 kcal/mol of soluble anhydrite up to $\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$ (Ref 18). The combination of the former value with the evaporation heat of water ($+0.51$ kcal/mol, Ref 14) gives the dissociation heat of selenate-plaster, i.e. 27.04 kcal. The value of the heat effect of an analogous process with sulfates determined by means of a differential-thermal analysis varies between 27.24 and 29.24 kcal/mol (Ref 18). The authors compute the formation heat of crystalline $\text{CaSeO}_4 \cdot 2\text{H}_2\text{O}$ from elements using the value of the hydration heat of CaSeO_4 (-6.02 kcal/mol) as well as the previously found value of crystalline CaSeO_4 from elements, and assuming that the formation heat of water is -68.32 kcal/mol

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Physico-chemical Properties of Selenates. V. Formation SOV/153-2-4-1/32
Heat of Calcium Selenate From Elements

(Ref 14): $\Delta H_{298.16}^{\circ}$ = -270.11±0.40 kcal/mol. The difference between the formation heats of CaSO_4 and CaSeO_4 is 70.16 kcal/mol. The energy of the lattice (U) of the low-temperature modification of CaSeSO_4 could be computed according to the known equation on account of the results obtained. Thus U of the CaSeO_4 modification mentioned last is -609.7 kcal/mol as determined by the authors. Similarly, the value of -646.8 kcal/mol for soluble anhydrite is found. A. F. Kapustinskiy, Corresponding Member AS USSR, contributed to the paper by expressing his criticism. In addition, M. S. Stakhanova is mentioned in the paper. There are 3 tables and 20 references, 8 of which are Soviet.

ASSOCIATION: Moskovskiy khimiko-tehnologicheskiy institut imeni D. I. Mendeleyeva, Kafedra obshchey i neorganicheskoy khimii (Moscow Institute of Chemical Technology imeni D. I. Mendeleev, Chair of General and Inorganic Chemistry)

SUBMITTED: May 12, 1958

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5(4)
AUTHORS:

Selivanova, N. M., Shneyder, V. A.

67031
SOV/153-2-5-1/31

TITLE:

Physico-chemical Properties of the Selenates. VI. Solubility
of Calcium Selenate in Water at Different Temperatures

PERIODICAL:

Izvestiya vysshikh uchebnykh zavedeniy. Khimiya i
khimicheskaya tekhnologiya, 1959, Vol 2, Nr 5, pp 651-656 (USSR)

ABSTRACT:

The authors studied the solubility of calcium selenate
starting from the cryohydrate point to the boiling point of
the saturated solution. The results are given in table 1.
The chemical analysis as well as the microscopic and
roentgen examination showed that the composition of the ground
phase, i.e. $\text{CaSeO}_4 \cdot 2\text{H}_2\text{O}$ is constant between 0 and 101.8°C. The
calcium content in this phase (obtained at 101.8°) is 18.21%,
the selenium content is 35.80%. The refractive indices are:
 $N_g = 1.576$, $N_p = 1.563$, $d_4^{20} = 2.68$, which were determined by
A. A. Mayer. Table 2 proves that the interplanar spacings,
calculated from the roentgenogram according to Debye-Scherrer,
are in accordance with the values of the previously synthesized
crystalline $\text{CaSeO}_4 \cdot 2\text{H}_2\text{O}$ and with the interplanar spacings of ✓

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SOV/153-2-5-1/31

Physico-chemical Properties of the Selenates. VI.
Solubility of Calcium Selenate in Water at Different Temperatures

the ground phase obtained by mixing the $\text{CaSeO}_4 \cdot 2\text{H}_2\text{O}$ in water at 0°C. Table 1 shows that the solubility of the calcium selenate constantly decreases with increasing temperatures. By comparison with the solubility of gypsum (Ref 6) the higher degree of solubility of the selenate, as compared to the sulfates (Ref 7), is again confirmed. The authors have previously established (Refs 8, 9) that the temperature coefficients of the solubility of SrSeO_4 and SrSO_4 on one hand, and that of BaSeO_4 and BaSO_4 on the other, have opposite signs, that of the selenates being negative. Therefore the law of S. S. Chin (Ref 10) did not prove true for the selenates and sulfates of calcium, strontium and barium. The authors did not succeed in determining the solubility of anhydrous calcium selenate in water. The data obtained by the authors on the solubility of $\text{CaSeO}_4 \cdot 2\text{H}_2\text{O}$ was used for the calculation of the solution-temperature in water, which is 1.820 kcal/mol. From this it was possible to calculate the value of the integral solution-temperature according to the

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Physico-chemical Properties of the Selenates. VI. SOV/153-2-5-1/31
Solubility of Calcium Selenate in Water at Different Temperatures

following formula, with the preparation of a saturated solution
at 25°C (Q_{satur}): $H = C_{\text{satur}} \frac{Q_{\text{satur}} - Q_{\text{unsatur}}}{C_{\text{satur}} - C_{\text{unsatur}}} + Q_{\text{satur}}$, in
which $Q_{\text{unsatur}} = 2.031$ kcal and C is a concentration of 1 mol
of saturated solution in 100/C mol of the solution agent. For
additional calculations they used the thermodynamic data for
calcium ions, for selenate ions and for water (Ref 21). The
data concerning the $\text{CaSeO}_4 \cdot 2\text{H}_2\text{O}$ make possible the calculation
of the entropy. From its change in the reaction (1) and, with
consideration of the entropy-values of selenium, oxygen,
hydrogen and metallic calcium (Refs 21, 22), the absolute
entropy of crystalline $\text{CaSeO}_4 \cdot 2\text{H}_2\text{O}$ is calculated. Professor

A. F. Kapustinskiy, Corresponding Member of the AS U:SR, is
mentioned in the article. The calcium selenate was prepared in
the "Krasnyy Khimik" ("Red Chemist") Works. There are
3 tables and 23 references, 12 of which are Soviet.

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67031

Physico-chemical Properties of the Selenates. VI. SOV/153-2-5-1/31
Solubility of Calcium Selenate in Water at Different Temperatures

ASSOCIATION: Moskovskiy khimiko-tehnologicheskiy institut imeni D. I.
Mendeleyeva; Kafedra obshchey i neorganicheskoy khimii
(Moscow Chemical-technological Institute imeni D. I.
Mendeleyev, Chair of General and Inorganic Chemistry)

SUBMITTED: May 12, 1958

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5(2)
AUTHORS:

Selivanova, N. M., Shneyder, V. A., Strel'tsov, I. S. SOV/78-4-7-3/44

TITLE: The Thermal Decomposition of Calcium Selenate (Termicheskoye
ra zlozheniye selenata kal'tsiya)

PERIODICAL: Zhurnal neorganicheskoy khimii, 1959, Vol 4, Nr 7,
pp 1481-1487 (USSR)

ABSTRACT: The heating- and cooling-curves of calcium selenate (Fig 1)
were plotted by means of the N. S. Kurnakov-pyrometer. The
temperature at the beginning of decomposition and at complete
decomposition, the melting temperature, and the degree of ther-
mal decomposition between 200-1150° were determined. The salt
 $\text{CaSeO}_4 \cdot 2\text{H}_2\text{O}$ served as a starting basis. The heating curve up
to 600° develops in a manner similar to that in the case of
gypsum. The cooling curve is not in agreement with the heating
curve, because irreversible processes occur in the case of
heating. The data of the analysis of the salt annealed at dif-
ferent temperatures are given by table 1. Table 2 mentions the
losses in weight at various temperatures. A microscopical in-
vestigation of the salt heated up to 200° shows a mixture of

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The Thermal Decomposition of Calcium Selenate

SOV/78-4-7-3/44

$\text{CaSeO}_4 \cdot 2\text{H}_2\text{O}$ -crystals and fine needle-shaped crystals of the non-aqueous selenate. Attempts made to produce the semihydrate $\text{CaSeO}_4 \cdot \frac{1}{2}\text{H}_2\text{O}$ were unsuccessful. During heating, modification changes sometimes occur. At 698° decomposition and conversion into selenite begins: $2\text{CaSeO}_4 \rightarrow 2\text{CaSeO}_3 + \text{O}_2$ with partial volatilization. The final product is calcium oxide and anhydride of the selenic acid. The results obtained do not explain the reversible endothermal effect at 786° of the heating curve. The results of the Debye-Scherrer X-ray pictures are given by table 3. As shown by table 4, calcium selenate is less temperature-resistant than calcium sulfate. The refraction indices of the crystals were determined by A. I. Mayer. There are 1 figure, 4 tables, and 23 references, 10 of which are Soviet.

SUBMITTED: April 11, 1958

Card 2/2

SELIVANOVA, N.M.; SHMEYDER, V.A.; STREL'TSOV, I.S.

Physicochemical study of selenates. Part 9: Thermal decomposition of magnesium selenate. Izv. vys. uchet. zav; khim. i khim. tekhn. 3 no. 5:787-793 '60.

(MIRA 13:12)

1. Moskovskiy khimiko-tehnologicheskiy institut imeni D.I.Mendeleyeva. Kafedra obshchey i neorganicheskoy khimii.
(Magnesium selenate)

3/063/60/005/003/003/011/XX
A051/A029

AUTHORS: Selivanova, N.M., Shneyder, V.A.

TITLE: On the Thermal Decomposition of Magnesium Selenate

PERIODICAL: Zhurnal Vsesoyuznogo Khimicheskogo Obshchestva im. D.I. Mendeleyeva, 1960, Vol. 5, No. 3, pp. 353-354

TEXT: A study was made on $MgSeO_4 \cdot 6H_2O$ usually stable under normal conditions. According to the experimental analysis, the content of magnesium and selenium in the initial salt was 8.59 and 28.78%, compared to the theoretical values of 8.83 and 28.68%; $d_4^{20} = 2.01$; $N_g = 1.495$; $N_p = 1.464$. The behavior of this salt at high temperatures was investigated by recording heating and cooling curves with a Kurnakov pyrometer. The experimental procedure was as follows: the salt was placed in a corundum crucible, which was inserted into a ceramic block. The heating rate of the furnace was $8-10^{\circ}C$ per min. Magnesium oxide was used as the inert substance. The thermocouple was platinum-platino-rhodium. Fig. 1 is the curve of the heated product. The cooling

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S/063/60/005/003/009/011/XX
A051/A029

On the Thermal Decomposition of Magnesium Selenate

curve of the magnesium selenate does not correspond to its heating curve, i.e., the processes which take place during heating of this salt under experimental conditions are non-reversible. Fig 2. is the curve of the anhydrous magnesium selenate plotted under the same conditions. It shows no thermal effects up to 400°C. All the thermal effects below the indicated temperature are determined, therefore, by the dehydration of this salt and by phenomena connected with this process. A study was made of the change in weight and composition of magnesium selenate in order to establish the causes of the existing thermal effects on the obtained curves (Fig.1,2), i.e., of hexahydrate and anhydrous salt during heating. The salts were calcinated in a crucible oven, into which an open scurundum crucible was placed, containing the material investigated; the calcination lasted for 1 hour without admission of air into the working space of the furnace. Fig. 3 shows the change in composition of $MgSeO_4 \cdot 6H_2O$ during the calcination. At 100°C the weight loss represented 13.5%,⁴ as against the theoretical value of 13.1% in the transformation: $MgSeO_4 \cdot 6H_2O \rightarrow MgSeO_4 \cdot 4H_2O$ (1). The data of the chemical

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A051/A029

On the Thermal Decomposition of Magnesium Selenate

analysis confirm this transformation. The Debyogram of the product obtained at 80°C is similar to the roentgenogram of previously synthesized pure $MgSeO_4 \cdot 4H_2O$; the specific gravity $d^{20} = 2.21$ corresponds to the same value of synthesized $MgSeO_4 \cdot 4H_2O$. Thus, the first endothermal effect (at 60°C) on the heating curve is explained by the transformation (1). The high endothermal effect at 118°C on the heating curve is explained by the further dehydration to a mixture of tetra- and monohydrates, which is accompanied by a partial reduction $Se^{6+} \rightarrow Se^{4+}$ and the partial volatilization of selenium. The effect at 260-230°C corresponds to the complete dehydration of magnesium selenate. According to the analysis the content of magnesium in the products of calcination of $MgSeO_4 \cdot 6H_2O$ at 300 and 400°C is 15.50 and 16.55% as against the theoretical value of 14.52% for $MgSeO_4$. Lines characteristic of $MgSeO_4$ are predominant on the Debyograms of these products. The endothermal effect at 170°C is explained by the formation of the intermediary product of dehydration, the composition of which was not established. The slight exothermal effect at 450°C is determined by the oxidation of Se^{4+} to Se^{6+} , and

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A051/A029

On the Thermal Decomposition of Magnesium Selenate

the effect at 620°C is caused by the partial decomposition of $MgSeO_4$ with the formation of magnesium selenate. The endothermal effects on the heating curves of $MgSeO_4 \cdot 6H_2O$ and $MgSeO_4$ at 760 and 770°C correspond to the complete decomposition of magnesium selenate to magnesium oxide. The weight loss of the salts at these temperatures is about 86.03% for $MgSeO_4 \cdot 6H_2O$ and 75.68% for $MgSeO_4$, as against the theoretical values of 85.94 and 76.18%, respectively, for the transformation:

$MgSeO_4 \cdot 6H_2O \rightarrow MgO$ and $MgSeO_4 \cdot 4H_2O \rightarrow MgO$, respectively. The data of a chemical analysis and X-ray studies confirm this transformation. The data in the table show that the liberation of the first molecules of water from magnesium selenate takes place earlier than in sulfate, but the last molecule in both salts is separated with difficulty. For magnesium sulfate this is explained by the different bond of this molecule of water with the salt (Refs. 1,4) or by its different position in the crystal lattice (Refs. 5-7). The same explanation is offered for magnesium selenate salt due to the iso-

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A051/A029

On the Thermal Decomposition of Magnesium Selenate

morphism of these salts. Comparison of the melting points of the two salts shows that the selenate is thermally less stable. This fact was established formerly (Ref 8,9). There are 3 graphs, 1 table and 9 references: 7 Soviet, 2 German.

SUBMITTED: April 16, 1959

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S/063/60/005/003/009/011/XX
A051/A029

On the Thermal Decomposition of Magnesium Selenate

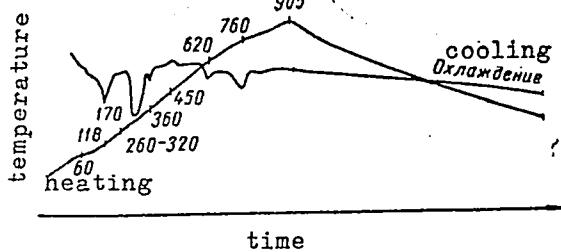


Figure 1:

Heating and cooling curves of magnesium selenate hexahydrate.

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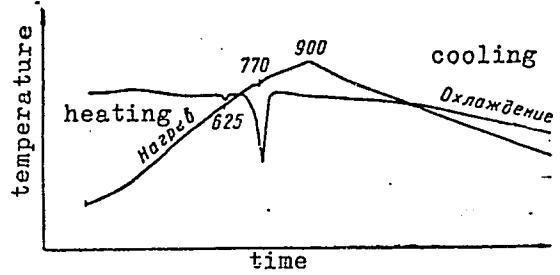


Figure 2:

Heating and cooling curves of anhydrous magnesium selenate.

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A051/A029

On the Thermal Decomposition of Magnesium Selenate

Table 1: Dehydration and decomposition temperature of magnesium selenate and sulfate.

Magnesium selenate transformation	temp. in °C	magnesium sulfate transformation	temp. in °C
$\text{MgSO}_4 \cdot 6\text{H}_2\text{O} \rightarrow$ $\leftarrow \text{MgSO}_4 \cdot 4\text{H}_2\text{O}$	60	$\text{MgSO}_4 \cdot 6\text{H}_2\text{O} \rightarrow$ $\leftarrow \text{MgSO}_4 \cdot 4\text{H}_2\text{O}$	90-100
$\text{MgSeO}_4 \cdot x\text{H}_2\text{O} \rightarrow$ $\leftarrow \text{MgSeO}_4$	300-320	$\text{MgSO}_4 \cdot x\text{H}_2\text{O} \rightarrow$ $\rightarrow \text{MgSO}_4$	300-340 ¹
beginning of decomposition: $\text{MgSeO}_4 \rightarrow$ $\leftarrow \text{MgSeO}_3 + \frac{1}{2}\text{O}_2$	620	$\text{MgSO}_4 \rightarrow \text{MgO} + \text{SO}_3$	1085 ³
complete decomposition: $\text{MgSeO}_3 \rightarrow \text{MgO} + \text{SeO}_2$ $\text{MgSeO}_4 \rightarrow \text{MgO} + \text{SeO}_3$	760	complete decomposition: $\text{MgSO}_4 \rightarrow \text{MgO} + \text{SO}_3$	1140 ³

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85607

S/078/60/005/010/027/030/XX
B017/B067

5 2640 2209, 1273, 1043

AUTHORS: Selivanova, N. M., Shneyder, V. A., and Strel'tsov, I. S.TITLE: Production of Crystal Hydrates of Magnesium Selenate ✓PERIODICAL: Zhurnal neorganicheskoy khimii, 1960, Vol. 5, No. 10,
pp. 2269-2271

TEXT: The crystal hydrates of magnesium selenate were synthesized. The compound $MgSeO_4 \cdot 6H_2O$ was produced by dissolving a stoichiometric amount of purest magnesium oxide in selenic acid, and subsequent crystallization at room temperature. Its specific gravity is 2.01; $MgSeO_4 \cdot 6H_2O$ has the following refractive indices: $Ng = 1.495 \pm 0.002$, $Np = 1.464 \pm 0.002$. The interplanar spacings (d) and the relative lines of intensity of the X-ray pictures of the following compounds are summarized in a table: $MgSeO_4 \cdot 6H_2O$, $MgSeO_4 \cdot 4H_2O$, $MgSeO_4 \cdot H_2O$, and $MgSeO_4$. By heating the crystals of $MgSeO_4 \cdot 6H_2O$ in the air bath at temperatures of 40 and $70^{\circ}C$, a crystal hydrate $MgSeO_4 \cdot 4.5H_2O$ is formed. By heating the crystal hydrate

X

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5-2610 also 2308

84215
S/078/60/005/010/010/021
B004/B067

AUTHORS: Selivanova, N. M., Shneyder, V. A., Strel'tsov, I. S.

TITLE: Thermal Decomposition of Beryllium Selenate ✓

PERIODICAL: Zhurnal neorganicheskoy khimii, 1960, Vol. 5, No. 10,
pp. 2272-2279

TEXT: By means of the Kurnakov pyrometer the authors took the heating and cooling curves of beryllium-selenate tetrahydrate in the temperature range 50 - 800°C (Fig. 1). They compared them with the curves obtained for anhydrous beryllium selenate (Fig. 2). From these data as well as from the analyses of the residues on ignition of $\text{BeSeO}_4 \cdot 4\text{H}_2\text{O}$ (Table 1) and BeSeO_4 (Table 2), of the loss on ignition of these substances (Table 3), and the Debye-Scherrer X-ray picture (Table 4), they obtained the following results: At 75°C, $\text{BeSeO}_4 \cdot 4\text{H}_2\text{O}$ passes over into $\text{BeSeO}_4 \cdot 2\text{H}_2\text{O}$. At 146°C, a further water loss occurs; and at 213°C, anhydrous BeSeO_4 is formed. These processes are accompanied by a partial reduction of Se^{6+} to Se^{4+} and a

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"APPROVED FOR RELEASE: 08/23/2000

CIA-RDP86-00513R001549810020-1

SELIVANOVA, N.M.; SHNEYDER, V.A.; RYABOVA, R.I.

Heat of formation of magnesium selenate from the elements. Zhur.
neorg. khim. 6 no.1:27-33 '61. (MIRA 14:2)
(Magnesium selenate) (Heat of formation)

APPROVED FOR RELEASE: 08/23/2000

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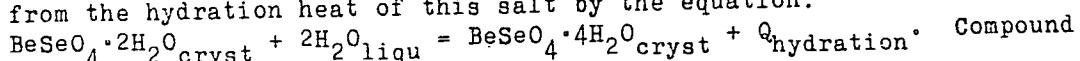
54750

1081, 1273, "60

S/076/61/035/003/009/023
B121/B203AUTHORS: Selivanova, N. M. and Shneyder, V. A.TITLE: Physicochemical properties of selenates. X. Heat of forma-
tion of beryllium selenate from elements

PERIODICAL: Zhurnal fizicheskoy khimii, v. 35, no. 3, 1961, 574-579

TEXT: The heat of reaction of beryllium selenate tetrahydrate with an aqueous solution of barium chloride at 25°C was determined with an isothermal calorimeter. The heats of formation of the crystalline beryllium salts $\text{BeSeO}_4 \cdot 4\text{H}_2\text{O}$, $\text{BeSeO}_4 \cdot 2\text{H}_2\text{O}$, and BeSeO_4 were also determined. $\text{BeSeO}_4 \cdot 4\text{H}_2\text{O}$ was supplied by the "Krasnyy khimik" plant and recrystallized twice before use. The correct composition of this salt was confirmed by chemical analysis. The heat of formation of $\text{BeSeO}_4 \cdot 4\text{H}_2\text{O}$ is: $\Delta H_f^{\circ} = -505.54$ kcal/mole. The heat of formation of crystalline $\text{BeSeO}_4 \cdot 2\text{H}_2\text{O}$ from elements was determined from the hydration heat of this salt by the equation:



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X

S/076/61/035/003/009/023

B121/B203

Physicochemical properties ...

D. Wagman, W. Evans, L. Levine, I. Iaffe, Selected values of chemical thermodynamic properties, Nat. Bur. Standards, 1952, Washington; N. V. Sidgwick, N. B. Lewis, J. Chem. Soc., London, 1290, 1926.

ASSOCIATION: Khimiko-tekhnologicheskiy institut im. D. I. Mendeleyeva
Moskva (Institute of Chemical Technology imeni D. I. Mendeleyev,
Moscow)

SUBMITTED: July 1, 1959

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X

SELIVANOVA, N.M.; SHNEYDER, V.A.; SAZYKINA, T.A.

Physicochemical properties of selenates. Part 12: Heat of formation of lithium selenates from elements. Izv.vys.ucheb.zav.;-khim.i khim.tekh. 5 no.2:183-187 '62. (MIRA 15:8)

1. Moskovskiy khimiko-tehnologicheskiy institut imeni D.I. Mendeleyeva kafedra obshchey i neorganicheskoy khimii.
(Lithium selenates) (Heat of formation)

REPUTUN, N.I.; AL'PEROVICH, D.I.; MIKHEYEV, V.N.; SHNEYDER, V.G.

Development of a method for expert testimony in alcoholic intoxication. Zhur. nevr. i psikh. 60 no.11:1523-1528 '60. (MIRA 14:5)

1. Kafedra sudebnoy meditsiny (zav. - prof. A.P.Kurdyumov) i psichiatrii (zav. - prof. D.S.Ozeretskovskiy) I Leningradskogo meditsinskogo instituta imeni I.P.Pavlova.
(ALCOHOLISM)

SHMEYDER, V.O.

D.A. Torba's design of a liquid pouring limiter. Transp. i khran.
nefti i neft'prod. no. 6122-23 '64. (MIRA 17:9)

1. Volgogradskaya proyektno-konstruktorskaya kontora Glavnogo upravleniya po transportu i snabzheniyu neft'yu i nefteproduktami RSFSR.

SHEMYDER, V.O.; KALTIKOV, V.M..

Automated service station. Transp. i zhran. nefti no.7:30-31 '63.
(MIRA 17:3)
1. Volgogradskaya proyektchnaya kontora Glavnogo upravleniya po trans-
portu i znabzheniyu neft'yu i nefteproduktami RSFSR.

USSR/Microbiology. Microbes Pathogenic for Man and
Animals

F

Abs Jour : Ref Zhur-Biol., No 13, 1958, 57744

Author : Shneyder V. P.

Inst : Not given

Title : Precipitation Reaction in Agar as a method of
Laboratory Diagnosis of Diphtheria

Orig Pub : Zdravookhr. Belorussii, 1957, No 3, 29-32

Abstract : The toxigenesis of diphtheria bacteria was de-
termined in Petri dishes according to the method
described by Ileik and Oukhterloni. 1.5 to 2.5%
MPA (pH 7.8) with 20 to 25% normal horse serum
was used as the culture medium. The diphtheria
antiseraum (Diaferm 3) mixed with agar and con-
taining 150, 50, 25, and 10 active units in ml
was poured into a groove preliminarily cut in

Card 1/3

USSR/Microbiology. Microbes Pathogenic for Man and F
Animals

Abs Jour : Ref Zhur-Biol., No 13, 1958, 57744

Abstract : units in 1 ml. An investigation of the microflora of the pharynx and nose of 133 diphtheria patients and persons suspected of having diphtheria, morphologically typical diphtheria bacteria were found in 61 cases and atypical in 8 cases. Positive precipitation reactions occurred in 6 cases in which positive results were obtained upon microscopic examinations of smears.

Card 3/3

PEN'KOV, A.I.; LOBKAREV, M.V.; SHNEYDER, V.P.

Effect of calcium muds on the stability of the well bore.
Burenje no.2:21-24 '65. (MIRA 12:5)

1. Turkmenskiy filial Vsesoyuznogo neftegazovogo nauchno-
issledovatel'skogo instituta.

ACC NR: AP6020678

SOURCE CODE: UR/0016/66/000/006/0026/0032

AUTHOR: Shneyder, V. P. (Member of infectious diseases dept)

ORG: Department of Infectious Diseases, Vitebsk Medical Institute (Kafedra infektsionnykh bolezney Vitebskogo meditsinskogo instituta)

TITLE: Passive Vi hemagglutination data on immunological shifts in typhoid patients treated with levomycetin and Vi antigen

SOURCE: Zh mikrobiol, epdiemiol i immunobiol, no. 6, 1966, 26-32

TOPIC TAGS: immunology, antibiotic, antibiotic therapy, ~~immunotherapy~~, typhoid, antigen, purified antigen, ~~Vi antigen~~, antibody, ~~titers~~, passive hemagglutination reaction, ~~BACTERIAL DISEASE~~

ABSTRACT:

A comparative study of changes in Vi-antibody titers in typhoid patients receiving different treatments was made. Group 1 was treated with purified Vi-antigen plus levomycetin and group 2 was given levomycetin alone. Generally, Vi-antibody titers rose in the group receiving Vi-antigen. Administration of the Vi-antigen is recommended for compound therapy, for controlling relapses, and for elimination of carriers. Orig. art. has: [W.A. 50; CBE No. 10] 2 tables and 2 figures.

SUB CODE: 06/ SUBM DATE: 25Dec65/ ORIG REF: 018/ OTH REF: 008/
Card 1/1 UDC: 616.927-085.372-059:615.779.931-07:616.927-097.3-078

SHNEYDER, V.V., referent

Central point supplying oxygen to converter plants (from "Iron and
Coal Trades Review," April 1957). Kislorod 11 no.6:37 '58.
(MIRA 12:3)

(France--Oxygen--Industrial applications)

SHAW VANCE, L.L.C.

John von Neumann
University of Princeton
Princeton, N.J.
1926

Baumgärtel, V. L., "Continuous functions on the closed plane and the Hahn-Banach theorem," *Duke Math. J.*, v. 19, p. 393-403, 1952.

This is a continuation of the first theorem, dealing with continuous functions. A set E is called σ -discrete if every open cover of E has a countable subcover. A compact Hausdorff space R is metrizable if and only if the diagonal $\{x = y : x, y \in R\}$ is a G_δ in $R \times R$. A compact Hausdorff space which is a continuous image of a Hausdorff space satisfying the second countability axiom is metrizable. A F_σ set in a topological space X is the σ -closure of S (cf. sets $\{S_i\}_{i=1}^\infty$ for example Hausdorff, *Measurable Functions*, 3rd ed., de Gruyter, 1933, pp. 90-93) obtained from the closed subsets of X . If a completely regular space E is an F_σ set in some compact extension of E , then it is an F_σ set in every compact extension. 4. A topological space X is said to have countable multiplicity if every family \mathfrak{M} of open subsets of X admits a countable subfamily \mathfrak{N} such that $\bigcup_{U \in \mathfrak{N}} U = \bigcup_{U \in \mathfrak{M}} U$. Let E be a Borel set in a space R of countable multiplicity, and let E' be a one-to-one continuous image of E lying in a compact Hausdorff space R' also of countable multiplicity. Then E' is a Borel set in R' .

E. Hewitt (Seattle, Wash.).

SHNEYDER, V. Ye.

Source: Matematicheskii Revue, 1950, no. 11, p. 101.

Schnieder, V. E. Descriptive theory of sets in topological

space. Uchenye Zapiski Moskov. Gos. Univ. 135. Mat-

ematika, Tom II, 37-85 (1948). (Russian)

This memoir extends a number of classical results concerning Borel sets and analytic sets from the case of subsets of a separable complete metric space to the case of subsets of more general topological spaces. The following conditions on the spaces X discussed are frequently employed: (A) every closed subset of X is a G_1 ; (B) if $\{G_\alpha\}$, $\alpha \in A$, is any family of open subsets of X , then there exists a countable subfamily $\{G_{\alpha_n}\}_{n=1}^\infty$ of $\{G_\alpha\}$, $\alpha \in A$, such that $\bigcup_{n=1}^\infty G_{\alpha_n} = \bigcup_{\alpha \in A} G_\alpha$. (The author describes spaces satisfying condition (B) as spaces of countable structure.) The Borel families constructed in the usual way by the operations of countable intersections and countable unions, starting with the families of open sets and closed sets, may be distinct, in general spaces, as may the corresponding families of analytic sets. A set analytic with respect to the family of closed sets is called an F_σ -set. The distinctions just described vanish, of course, for spaces satisfying condition (A).

A number of simple topological theorems are first proved, among them the facts that conditions (A) and (B) are equivalent for compact (bicomplete in the author's terminology) Hausdorff spaces and that every regular space satisfying condition (B) is completely normal. Next, the following space L is studied in detail. The space L is made up of two intervals $I = E[x; 0 \leq x < 1]$ and $I' = E[x'; 0 < x' \leq 1]$. A generic neighborhood of $x_0 \in I$ is $[x_0, x_0 + \delta] \cup (x_0 - \delta, x_0 + \delta)$; a generic neighborhood of $x_0 \in I'$ is $(x_0 - t, x_0] \cup [x_0 - t, x_0]$, where $x_0 = x_0'$ and t is an arbitrary positive real number. It is proved that L is a compact Hausdorff space satisfying condition (B) that is nonmetrizable and that no uncountable analytic subset of L can be represented as a continuous image of a subspace of the space of irrational numbers.

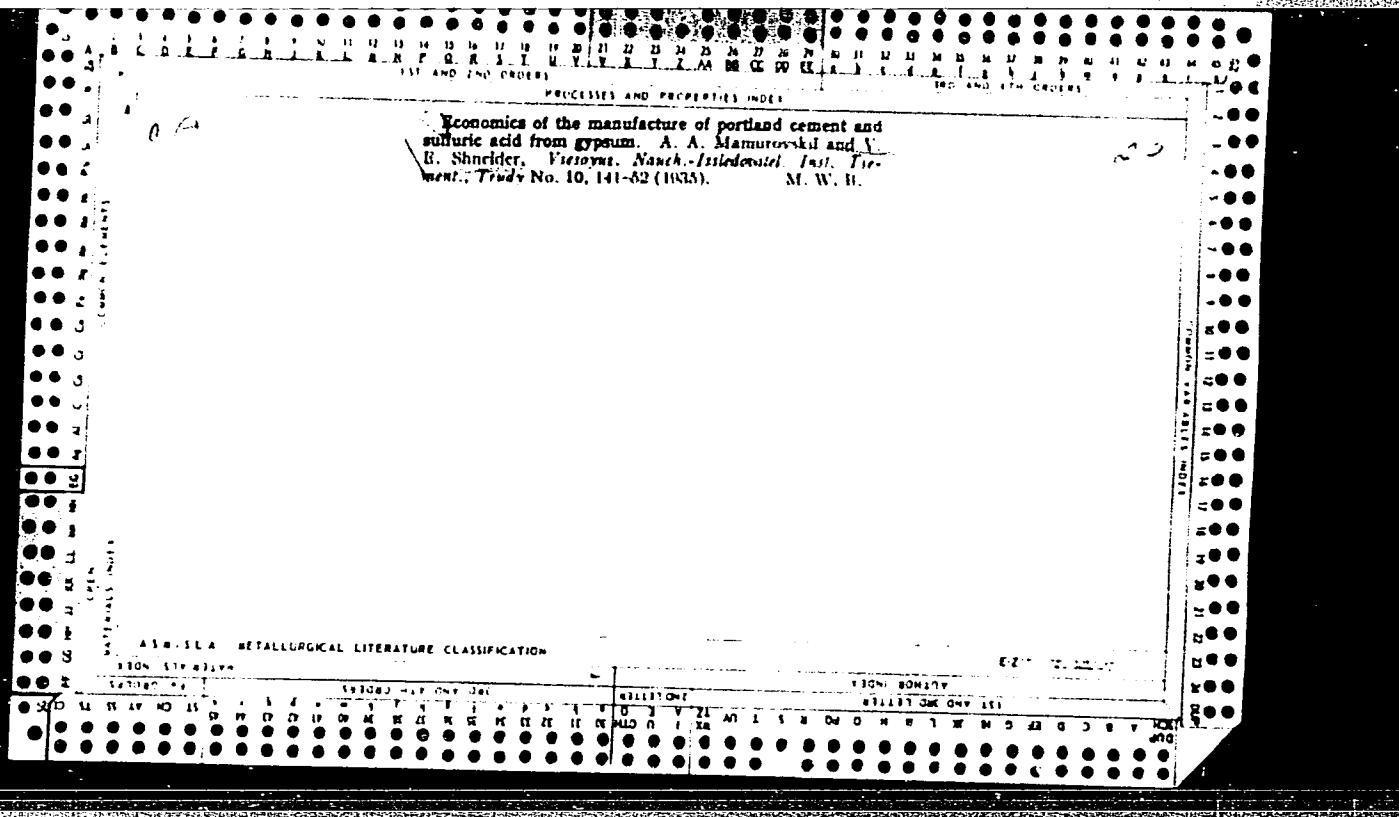
OCHAN, Yuriy Semenovich; SHNEYDER, Vladimir Yevgen'yevich; IGNAT'YEVA,
A.V., red.; SMIRNOVA, M.I., tekhn.red.

[Mathematical analysis; textbook for teachers' institutes] Matema-
ticheskii analiz; uchebnoe posobie dlia pedagogicheskikh institutov.
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(Mathematical analysis)

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ATC

Present status and prospects for development of sources
for Bauxite fluxes in the Urals. V. R. SARKISOV AND Yu.
I. CHERNOSVITOV. *Gornyi Zhurnal*, 1940, No. 12, pp.
28-33; *Khim. Referat. Zhur.*, 4 [5] 16 (1941). M.Ho.

SCHNEIDER, V. I.E.

Geography and Geology

Requirements of industry as to the quality of mineral raw materials. Handbook for geologists
Moskva, Gos. izd-vo geologicheskoi lit-ry Komiteta po delam geologii pri SNK SSSR,
No. 50, Gypsum, 1948

9. MONTHLY LIST OF RUSSIAN ACCESSIONS, Library of Congress, October 1952. Uncl.

GUSEV, Sergey Georgiyevich; VASIL'YEV, V.P., redaktor; SHNEYDER, V.Ye.,
kandidat ekonomicheskikh nauk, nauchnyy redaktor; SHPAYER, A.L.,
redaktor; PANOV, L.Ya., tekhnicheskiy redaktor.

[Accounting, bookkeeping and work analysis in enterprises of the
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BERKOVICH, T.M.; BOYAZNYY, L.S.; LUKOSHKINA, L.A.; DAVYDOVA, F.L.;
~~SHNEYDER, V.Y.~~, SHPAYER, A.L., redaktor; PYATAKOVA, N.D.,
tekhnicheskiy redaktor

[Manufacture of asbestos-cement elements] Proizvodstvo asbesto-
tsementnykh izdelii. Pod red. T.M.Berkovicha. Moskva, Gos.
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YERSHOV, A.D., glavnyy red.; NEKRASOVA, N.B., red.izd-va;
IVANOVA, A.G., tekhn.red.

[Industry's demands in the quality of mineral raw materials;
handbook for geologists] Trebovaniia promyslennosti k kachestvu
mineral'nogo syr'ia; spravochnik dlia geologov. Moskva, Gos.
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[Industry's requirements as to the quality of mineral raw
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izdat. No.52. [Cement raw materials] TSementnoe syr'e. 1962.
82 p. (MIRA 15:7)

1. Moscow. Vsesoyuznyy nauchno-issledovatel'skiy institut mine-
ral'nogo syr'ya.

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BERKOVICH, T.M., kand. tekhn. nauk; BOYAZNYY, L.S., inzh.; DAVYDOVA,
F.L., inzh.; LUKOSHINA, L.A., kand. tekhn.nauk; SHNEYDER,
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nauchnyy red.; TYUTYUNIK, M.S., red. izd-va; SHERSTNEVA, N.V.,
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[Manufacture of asbestos-cement products] Proizvodstvo asbesto-
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Berkovicha. Moskva, Gosstroizdat, 1962. 367 p.
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